Acidity Titration Method SM 18 th /20 th 2310 B						
Facility Name:	VELAP ID					
Assessor Name: Analyst Name	Analyst Name:		Inspection Date			
Relevant Aspect of Standards	Method Reference	Y	N	N/A	Comments	
Records Examined: SOP Number/ Revision/ Date Analyst:						
Sample ID: Date of Sample Preparation:		Date of Analysis:				
Were all stock and standard solutions and dilution water prepared with distilled or deionized water that has been freshly boiled for 15 minutes and cooled to room temperature?	3.a					
Was KHC ₈ H ₄ O ₄ crushed to a size of 100-mesh and dried at 120°C for 2 hours?	3.b					
Was 10.0±0.5g KHC ₈ H ₄ O ₄ diluted to 1000mL for a concentration of 0.05N?	3.b					
Was a sodium hydroxide titrant prepared to 0.1N and standardized by titrating 40.00 mL of the KHC ₈ H ₄ O ₄ solution to the inflection point of approximately 8.7pH?	3.c					
Was the normality of the sodium hydroxide titrant adjusted to 0.1N or measured with the measured value used in further calculations?	3.c					
Was a sodium hydroxide titrant prepared to 0.02N and standardized against 15.00mL KHC ₈ H ₄ O ₄ as above?	3.d					
If a sample was suspected to contain hydrolysable metals and reduced forms of polyvalent cations, was the sample's pH verified to be <4 and treated by adding 5 drops of hydrogen peroxide then boiled for 2 to 5 minutes?	4.a					
Were the samples adjusted to room temperature?	4.b					
If free residual chlorine was present in samples, was it neutralized by adding 1 drop of sodium thiosulfate to samples or exposing to UV light?	4.b					
Were the samples titrated over a white surface for adequate visualization of color change?	4.b					
When forming a potentiometric titration curve was the curve obtained smooth indicating equilibrium being reached between successive alkali additions?	4.c.2					
Notes/ Comments:						